Supplementary Information

Additive Manufacturing towards the Fabrication of Greener Electrochemical Sensors for Antioxidants

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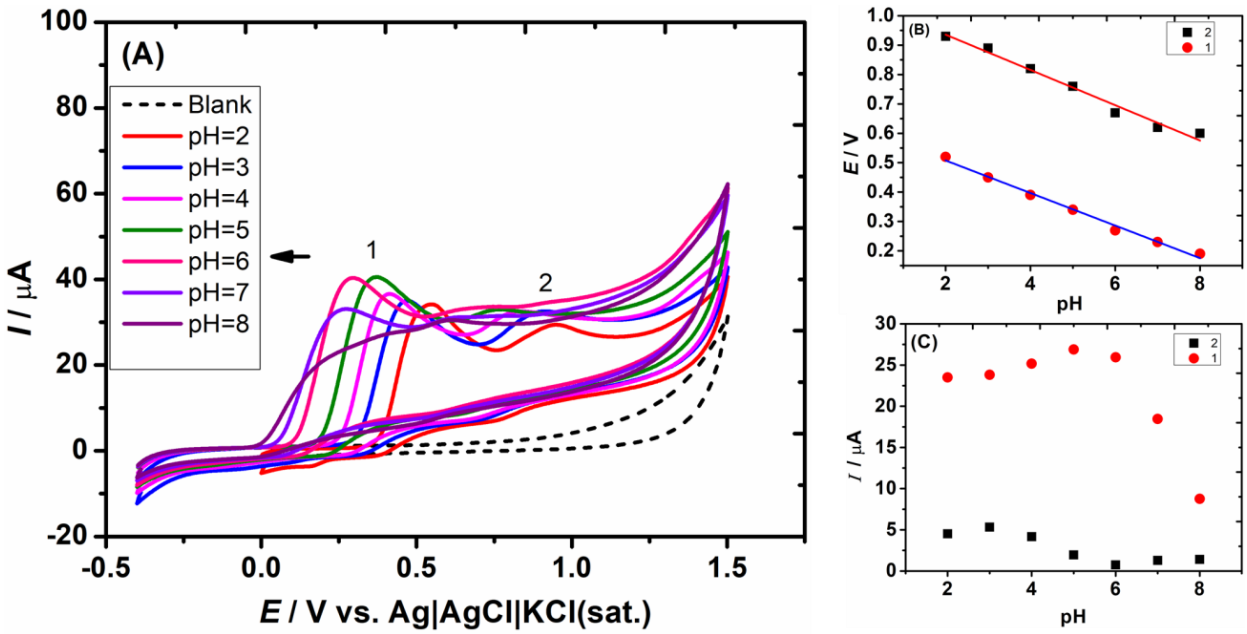

Figure S1. (A) Cyclic voltammetric recordings for 1 mmol L^{-1} PY in 0.12 mol L^{-1} BR buffer (pH range from 2.0 to 8.0). (B) pH influence at peak potential (Ep) and (C) pH influence at peak current (Ip).

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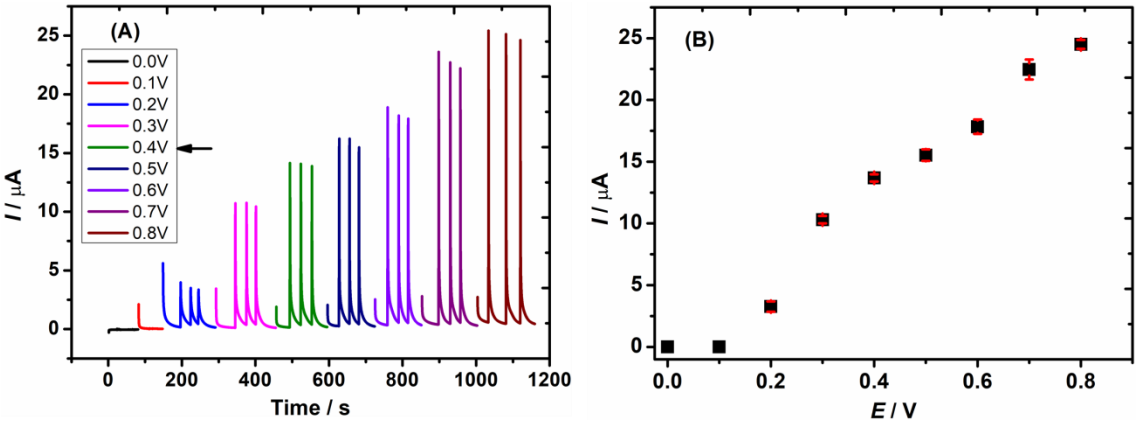

Figure S2. (A) BIA-AD recordings obtained from 3 successive injections of 100 µmol L^{-1} PY. (B) Hydrodynamic voltammogram obtained by plotting the peak current values (average of 3 injections) as function of the corresponding applied potential. Electrolyte: BR buffer (0.12 mol L^{-1}, pH 6.0). Conditions: injected volume: 100 µL and dispensing rate: 200 µL s^{-1}.

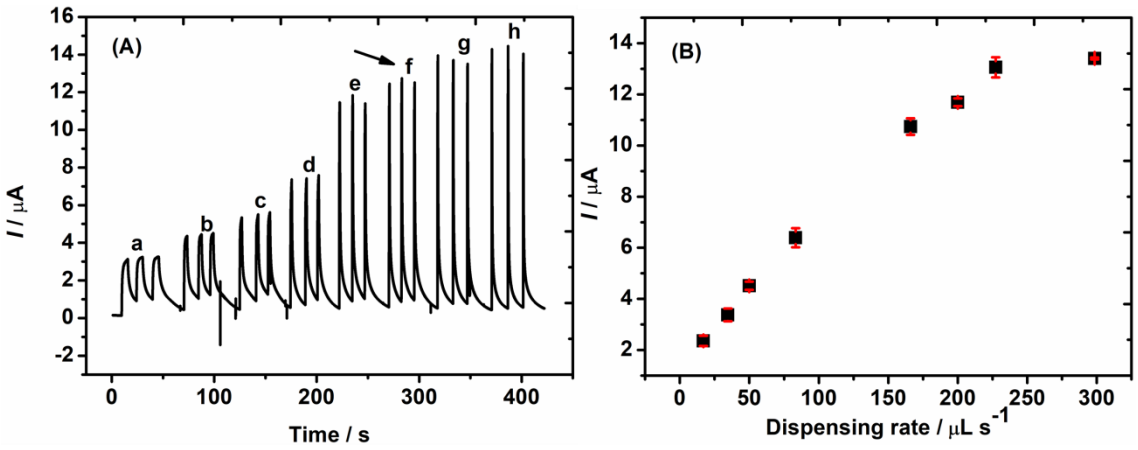

Figure S3. BIA-AD recordings obtained from 3 successive injections of 100 µmol L^{-1} PY as function of dispensing rate (a) 16.9, (b) 34.5, (c) 50, (d) 83.3, (e) 166, (f) 200, (g) 227.3 and (h) 300 µL s^{-1}. (B) Influence of the dispensing rate on peak current (Ip) of the PY. Conditions: injected volume: 100 µL and working potential: 0.4 V.

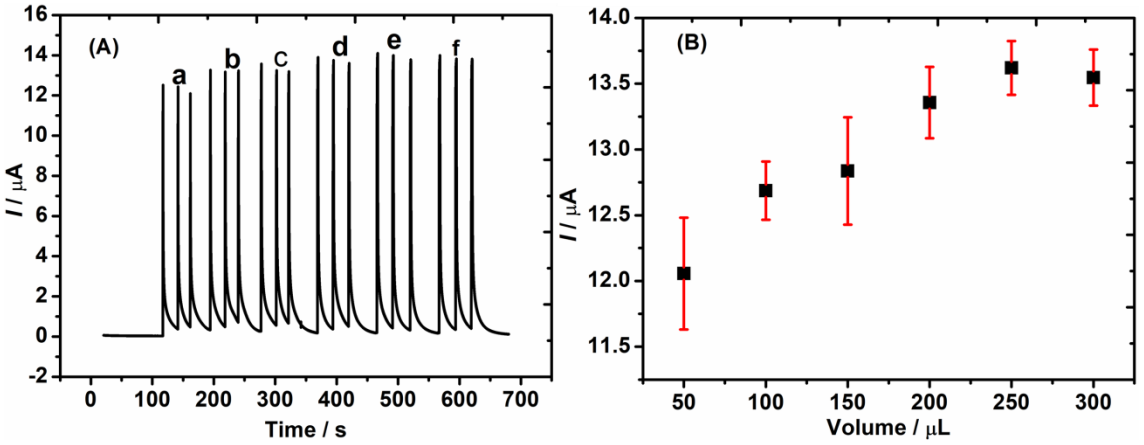

Figure S4. BIA-AD recordings obtained from 3 successive injections of 100 µmol L^{-1} PY as function of injected volume (a) 50, (b) 100, (c) 150, (d) 200, (e) 250 and 300 µL. (B) Influence of the injected volume on the peak current (Ip) of the PY. Conditions: dispensing rate: 200 µL s^{-1} and working potential: 0.4 V.